

Roll Number		
-------------	--	--

SET A



INDIAN SCHOOL MUSCAT
HALF YEARLY EXAMINATION
SUBJECT : CHEMISTRY

CLASS: XII

Sub.Code: 043

Time Allotted: 3 Hrs.

16.09.2019

Max.Marks:70

- **General Instructions:** All questions are compulsory.
- Marks for each question are indicated against it.
- Question numbers 1 to 20 are very short answer questions and carry 1 mark each.
- Question numbers 21 to 27 are short answer questions and carry 2 marks each.
- Question numbers 28 to 34 are also short answer questions and carry 3 marks each
- Question numbers 35 to 37 are long answer questions and carry 5 marks each.
- Use log tables if necessary, use of calculators is not allowed.

Pick out the correct option:

1. Phosgene is 1
 - a) Chloroform
 - b) Carbonyl chloride
 - c) Iodoform
 - d) Methylene chloride
2. The weakest intermolecular force is present in 1
 - a) Thermoplastics
 - b) Fibers
 - c) Elastomers
 - d) Thermosetting plastics
3. Among the following, which is classified as neurologically active drugs? 1
 - a) Antiseptic
 - b) Analgesic
 - c) Antifertility
 - d) Antacids
4. Artificial sweetener that has its use limited to cold food is 1
 - a) Saccharin
 - b) Sucralose
 - c) Alitame
 - d) Aspartame
5. The molal depression constant depends on 1
 - a) Nature of solvent
 - b) Nature of solute
 - c) Vapour pressure of solution
 - d) Number of particles formed in the solution
6. Which of the following solutions behave ideally? 1
 - a) Benzene and toluene
 - b) Ethanol and acetone

- c) Carbon disulphide and acetone
d) Phenol and aniline
7. To a Daniel cell, if an external voltage of 1.1V is applied 1
 a) Electrons flow from zinc to copper and current flows from copper to zinc
 b) Electrons flow from copper to zinc and current flows from zinc to copper
 c) No flow of electrons or current
 d) Zinc dissolves at anode and copper deposits at cathode
8. Which of the following metals is obtained by leaching the ore with dilute cyanide solution? 1
 a) Silver
 b) Titanium
 c) Nickel
 d) Aluminium
- 9 In an electrochemical cell 1
 a) Potential energy is converted to chemical energy
 b) Chemical energy is converted to potential energy
 c) Chemical energy is converted into electric energy
 d) Electrical energy is used to carry out a chemical reaction in the cell.
- 10 The following cannot be used to prevent corrosion 1
 a) Bisphenol
 b) Tin
 c) Zinc
 d) Carbon steel

Fill in the blanks:

- 11 Carbon attached to four different atoms/groups is called..... 1
- 12 -----is added to raw rubber (to form crosslinks) along with other appropriate additives to form vulcanized rubber. 1
- 13 -----detergent has germicidal properties. 1
- 14 Constant boiling mixtures are termed as..... 1
- 15 Purest form of iron is..... 1

Answer the following questions:

- 16 Arrange the following in order of increasing reactivity in nucleophilic substitution reaction: 1
 CH_3F , CH_3I , CH_3Br , CH_3Cl
- 17 Draw the structure of 4-methoxy acetophenone. 1
- 18 What are the main constituents of Dettol? 1
- 19 How is K_H of a gas in a liquid related to its solubility? 1
- 20 Give one advantage of $\text{H}_2\text{-O}_2$ fuel cell. 1
- 21 Convert 2
 a) Benzene to biphenyl
 b) 2-Bromopropane to 1-Bromopropane

- 22 Give reason 2
- a) Vinyl halides are inert to nucleophilic substitution reaction
- b) (dl)-butan-2-ol is optically inactive
- 23 a) Give a chemical test to distinguish between 2
- 2-Methylpropanol and 2-Methyl propan-2-ol.
- b) Write the IUPAC name of o-cresol.
- 24 A solution was formed by mixing chloroform with acetone. 2
- a) Identify the type of deviation shown by this solution? Give reason for your choice.
- b) Give one characteristic of this type of solution.
- OR**
- i) Define reverse osmosis.
- ii) What happens when red blood cells are placed in 0.5% sodium chloride solution?
- 25 Write the reaction taking place at the anode and cathode of lead storage battery. 2
- OR**
- a) What is the role of salt bridge?
- b) Predict the product of electrolysis of aqueous solution of silver nitrate using silver electrodes.
- 26 What is the role of 2
- a) Sodium cyanide in the froth floatation process
- b) Calcium oxide in the extraction of iron from its ore?
- 27 What are biodegradable polymers? Name the monomers of a biodegradable polyamide. 2
- 28 a) What are enantiomers? 3
- Draw the structures of the possible enantiomers of 3-Methylpent-1-ene.
- b) State Saytzeff rule.
- 29 Write the mechanism of dehydration of ethanol at 443K. 3
- 30 Write the structure of repeating monomeric unit of 3
- a) Terylene
- b) Neoprene
- c) Teflon
- 31 Mention the action of the following on human body in bringing relief from a disease 3
- a) Brompheniramine
- b) Aspirin
- c) Equanil

OR

Explain the following with an example

- i. Antacids
- ii. Disinfectant
- iii. Narrow spectrum antibiotics

- 32 How many mL of 0.1M HCl is required to react completely with 1g mixture of sodium carbonate and sodium hydrogen carbonate containing equimolar amounts of the two? 3

OR

Calculate the mass of a non-volatile solute (molar mass 40g/mol) which should be dissolved in 114g octane (molar mass 114 g/mol) to reduce the vapour pressure to 80%.

- 33 Determine the value of equilibrium constant and Gibb's energy for the following reaction 3
- $$\text{Ni}_{(s)} + 2\text{Ag}^+_{(aq)} \rightarrow \text{Ni}^{2+}_{(aq)} + 2\text{Ag}_{(s)}$$
- $E^0 = 1.05\text{V}$, $1F = 96500\text{C/mol}$.

OR

In a copper-silver cell, the concentration of copper ions is 0.10M and the concentration of silver ions is not known. The cell potential when measured was 0.422V. Determine the concentration of silver ions in the cell. [Given $E^0_{\text{Ag}^+/\text{Ag}} = +0.80\text{V}$, $E^0_{\text{Cu}^{2+}/\text{Cu}} = +0.34\text{V}$]

- 34 Explain 3
- a) Mond's process
 - b) Zone refining
 - c) Column chromatography for purification of rare elements.
- 35 Illustrate 5
- i) Williamsons synthesis
 - ii) Reimer Tiemann reaction
- b) What happens when phenol is treated with- [write chemical equations]
- i) Bromine water
 - ii) acidified sodium dichromate
 - iii) Zinc dust

OR

An organic compound A [C_6H_6O] gives a characteristic color with neutral ferric chloride. When carbon dioxide gas is passed through an alkaline solution of A at 400K under pressure compound B is obtained. Compound B on acidification gives C which reacts with acetyl chloride to form D which is a popular pain killer. Deduce structures of A, B, C and D and write the reaction of conversion of C to D.

- 36 i. Why a person suffering from high blood pressure is advised to take minimum quantity of common salt? 5
- ii. Define the term Colligative properties
- iii. A solution of glycerol [molar mass 92] in water was prepared by dissolving some glycerol in 500g of water. This solution has a boiling point of $100.42^{\circ}C$. What mass of glycerol was dissolved to make the solution? [Given: Boiling point of pure water= $100^{\circ}C$, K_b for water= $0.512K$ kg/mol]

OR

- a) Why do scuba divers carry oxygen cylinders diluted with helium?
 - b) Will the value of Van't Hoff factor be >1 or <1 , when a solute undergoes dissociation?
 - c) An aqueous solution containing 12.48g of $BaCl_2$ in 1kg of water boils at 373.08K. Calculate the degree of dissociation of $BaCl_2$. [molar mass of $BaCl_2=208.34g/mol$, K_b of water= $0.52 K$ kg/mol]
- 37 i. State Kohlrausch's law of independent migration of ions. 5
- ii. Mention one application of Kohlrausch's law.
- iii. The electrical resistance of a column of 0.05M NaOH solution of 1cm diameter and length 50cm is 5.55×10^3 ohm. Calculate its resistivity, conductivity and molar conductivity.

OR

- a) Why does conductivity of solution decrease with dilution?
- b) State Faraday's first law of electrolysis.
- c) A voltaic cell is set up at $25^{\circ}C$ with the following half cells, Al^{3+} [0.001M] and Ni^{2+} [0.10M]. Write an equation for the reaction that occurs when cell generates an electric current and determine the cell potential. [Given: $E^{\circ} Ni^{2+}/Ni = -0.25V$, $E^{\circ} Al^{3+}/Al = -1.66V$]

End of the Question Paper